. 1. Write the equilibrium constant expression for this reaction. _____

   The following changes are introduced at time T_f. Determine which situation described in the box that follows best describes how the graph will be changed to the right of T_f. Write the letter of your choice in the space provided and describe how Le Châtelier's principle predicts the result you have chosen.

2. Temperature of the system is increased _____

   _____

3. Pressure of the system is increased _____

   _____

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a. The HI curve rises, the I_2 curve drops, the H_2 curve remains the same.

b. The HI curve drops, the I_2 curve rises, the H_2 curve drops.

c. The HI curve rises, the I_2 curve rises, the H_2 curve drops.

d. The curves for all three components rise.

e. The H_2 and I_2 curves rise, the HI curve drops.

f. The H_2 and I_2 curves drop, the HI curve rises.

g. After an initial rise, the I_2 curve drops, the H_2 curve drops, the HI curve rises.

h. All curves remain the same.

Critical Thinking

Make a concept map beginning with the term **equilibrium constant expression**. Complete the map with terms from Chapter 18 of your text.